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Fatiha Abidar, Mohamed Morghi, M'hamed Abali, Fouad Sinan, Mohamed Chiban, Hassan Eddaoudi and Mohamed Zebret

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### Article abstract

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# ORTHOPHOSPHATE ION ADSORPTION ONTO RAW SHRIMP SHELLS

*Adsorption des ions orthophosphates sur carapaces de crevettes à l'état brut*

Fatima ABIDAR, Mohamed MORGHI, M'bamed ABALI, Fouad SINAN, Mohamed CHIBAN, Hassan EDDAOUDI, Mohamed ZERBET\*

Université Ibn Zohr, Faculté des Sciences, BP 8106 Hay Dakhla, Agadir, Morocco

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## ABSTRACT

The aim of this work was to develop new low-cost adsorbents obtained from animal origins, available in large quantities and environmentally friendly. Raw shrimp shell (RSS), a biomaterial of animal origin, is abundant, available, renewable and non-toxic. It has physicochemical properties that can induce a significant adsorptive activity. In this study, the removal of phosphate anions ( $\text{H}_2\text{PO}_4^-$ ,  $\text{HPO}_4^{2-}$ ) from aqueous solution by adsorption onto raw shrimp shells was studied. The surface micro-morphology of the biomaterial was investigated using scanning electron microscope and qualitative element composition was analyzed using energy dispersive X-ray and infrared spectroscopies. The efficiency of the biomaterial was investigated using a batch adsorption technique under different experiment conditions, achieved by varying parameters such as adsorbent dosage, the contact time, the initial phosphate anion concentrations, the temperature and the initial solution pH. Results show that the kinetics adsorption of phosphate ions by the biomaterial is relatively quick and the biomaterial showed a high adsorption capacity of  $0.20 \text{ g}\cdot\text{g}^{-1}$  and  $0.4 \text{ g}\cdot\text{g}^{-1}$  for  $\text{HPO}_4^{2-}$  and  $\text{H}_2\text{PO}_4^-$ , respectively. The adsorption data were analyzed using the Langmuir, Freundlich and Temkin adsorption isotherms to

determine the nature of the adsorption sites. Both Langmuir and Freundlich adsorption models showed good fits to the experimental adsorption data.

**Key Words:** *Adsorption, orthophosphate, shrimp shell, wastewater.*

## RÉSUMÉ

Cette étude s'inscrit dans le cadre de certains de nos travaux visant la valorisation de matériaux naturels d'origine végétale, animale et minérale pouvant rendre les procédés d'épuration des eaux simples et moins coûteux. Les carapaces de crevettes à l'état brut, biomatériau d'origine animale, entrent dans cette catégorie. Il est abondant, disponible, renouvelable et non toxique, et présente des propriétés physicochimiques qui peuvent induire une activité adsorbante importante. Le biomatériau a été caractérisé par microscopie électronique à balayage, par spectroscopie d'émission X et spectroscopie infrarouge. Une étude complète de l'adsorption des ions phosphates ( $\text{HPO}_4^{2-}$  et  $\text{H}_2\text{PO}_4^-$ ) sur des carapace de

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\*Corresponding author:  
Phone: + 212 (0) 5 28 22 09 57  
Email: zerbetmohamed@gmail.com

crevettes à l'état brut a été effectuée. L'influence de paramètres physicochimiques tels que la masse d'adsorbant, le temps de contact, le pH, la température et la concentration initiale de phosphate a été étudiée. Les résultats obtenus montrent une cinétique rapide et une grande capacité de ce biomatériau à retenir les ions phosphates pouvant atteindre  $0,20 \text{ g}\cdot\text{g}^{-1}$  pour  $\text{HPO}_4^{2-}$  et  $0,40 \text{ g}\cdot\text{g}^{-1}$  pour  $\text{H}_2\text{PO}_4^-$ . Les isothermes d'adsorption étudiés (Langmuir, Freundlich et Temkin) montrent une bonne corrélation avec les modèles de Langmuir et de Freundlich.

**Mots-clés :** Adsorption, orthophosphates, carapace de crevette, eaux usées.

## 1. INTRODUCTION

Phosphorus is the key nutrients for biological and chemical processes in natural water bodies, and large quantity of phosphate present in wastewater is one of the main causes of eutrophication in surface waters (CHIBAN *et al.*, 2012; RODIER, 2009; NGUYEN *et al.*, 2012; VISSMAN and HAMMER, 2005). Morocco has very limited water resources. Besides drought cycles, the already scarce water is the subject of the continuous increasing needs and growing quality degradation due to pollution. The nature of the pollution charge of quality degradation is mainly organic type, nitrogen and phosphorus. The eutrophication of rivers and dams deductions was observed; oxygen levels in some reservoirs have fallen and have promoted specific anaerobic conditions. In terms of phosphorus, the values stored in some dams have exceeded  $0.5 \text{ mg}\cdot\text{L}^{-1}$  (OBSERVATOIRE NATIONAL DE L'ENVIRONNEMENT ET DU DÉVELOPPEMENT DURABLE DU MAROC, 2015).

Various techniques have been used for phosphate and heavy metals removal. Phosphate removal techniques fall into three main categories: physical, chemical and biological. The major disadvantage that we come across with the conventional processes is that the process is expensive and not eco-friendly (CHIBAN *et al.*, 2013; NGUYEN *et al.*, 2012; WAN NGAH and HANAFIAH, 2008). Though in the recent years the focus of the research is to use biocompatible materials as potential adsorbents. This solution can prove unpollutant, economic and able to bring cost effectiveness. Research based on animal, vegetal and inorganic adsorption materials, have been used by many researchers (CHUBAR *et al.*, 2005; DE LIMA *et al.*, 2012; DIVYA JYOTHI *et al.*, 2012; ISMAIL, 2012; KARACA *et al.*, 2006; LI *et al.*, 2006; NGUYEN *et al.*, 2012; SAAD *et al.*, 2007; WAN NGAH and HANAFIAH, 2008; YUCEF and ACHOUR, 2005; ZHANG *et al.*, 2010). An adsorbent can be considered as cheap or low-cost if it is abundant in nature, requires little processing and is a byproduct of waste

material from waste industry (BAILEY *et al.*, 1999). The raw shrimp shells, is a residue of animal origin, easily available and inexpensive, seems to be a suitable material for phosphate sorption. Actually chitin is one of the most abundant biopolymers in nature and is a major component of shrimp shells. Several studies have clearly demonstrated that chitin (and its deacetylated derivative, chitosan), have gained wide attention as effective biosorbents due to low cost and high contents of amino and hydroxyl functional groups which show significant adsorption potential for the removal of various aquatic pollutants (BHATNAGAR and SILLANPÄÄ, 2009). Moreover, the use of these shells as raw material for wastewater treatment could also increase the additional value of this biomass waste that could potentially become an environmental problem.

The present work was performed to evaluate the capacity of adsorption of the raw shrimp shells, for removing phosphate anions ( $\text{HPO}_4^{2-}$  and  $\text{H}_2\text{PO}_4^-$ ) from aqueous solutions. The study of the influence of various parameters such as adsorbent dosage, contact time, pH, temperature and initial adsorbate concentration on adsorption were investigated.

## 2. MATERIALS AND METHODS

### 2.1 Materials

The inert solid biomaterial (ISBM) used as support adsorption in this work is the raw shrimp shells (RSS). The shrimp shells were washed thoroughly with pure water several times. The washed materials were then air-dried ( $\approx 27^\circ\text{C}$ ), grinded and sieved ( $\leq 500 \mu\text{m}$ ). The microparticles of the raw shrimp shells are used as adsorbent materials in batch experiments without any pretreatment to avoid extra expenditure.

Crustacean shell waste consists mainly of 30-40% protein, 30-50% calcium-carbonate, and 20-30% chitin (JOHNSON and PENISTON, 1982; KURITA, 2006). Crab shells consist of  $\text{CaCO}_3$  and chitin, usually cross-linked with some protein, and a proportion of lipids (Figure 1) (JOHNSON and PENISTON, 1982; KURITA, 2006).

The point of zero charge (PZC) was determined using the solid addition method (BANERJEE and CHATTOPADHYAYA, 2013) *viz.*: 50 mL of 0.01 M NaCl solutions were taken in different conical flasks of 100 mL and 0.5 g of adsorbent was introduced in each of them. Now pH values of these solutions were adjusted in 2 to 12 range by 0.1 M HCl/NaOH solutions. These flasks were kept for 48 h and the final pH of the solutions was measured. Graphs were



Figure 1. Schematization of the raw shrimp shells (JOHNSON and PENISTON, 1982; KURITA, 2006).  
Schématisation de carapaces de crevettes à l'état brut (JOHNSON et PENISTON, 1982; KURITA, 2006).

plotted between  $\text{pH}_{\text{final}}$  versus  $\text{pH}_{\text{initial}}$ . The point of intersection of the curve of  $\text{pH}_{\text{final}}$  versus  $\text{pH}_{\text{initial}}$  was recorded as  $\text{pH}_{\text{PZC}}$  of the shrimp shells.

## 2.2 Solutions of the studied ions

$\text{NaH}_2\text{PO}_4$  and  $\text{Na}_2\text{HPO}_4 \cdot 12\text{H}_2\text{O}$  salts were analytical grade reagents from Fluka. Aqueous solutions of these salts were prepared in de-ionized water. The concentration range of the solution ion prepared varies between  $50 \text{ mg}\cdot\text{L}^{-1}$  and  $7 \text{ g}\cdot\text{L}^{-1}$ . For the influence of PH in the adsorption experiments, the pH of the solution was adjusted by adding dilute solution of hydrochloric acid (1 M HCl) or sodium hydroxide (1 M NaOH).

## 2.3 Batch adsorption studies

Batch adsorption experiments were carried out by batch process. 40 ml of phosphate solution with a given ions concentration ( $C_i$ ) ( $\text{mg}\cdot\text{L}^{-1}$ ), was mixed with 1 g of dried and grinded shrimp shells. The solutions put in contact with the raw shrimp shells matter were maintained at a constant temperature of  $25^\circ\text{C}$  in a water bath thermostat, the mixture being vigorously stirred at 180 rpm. The sampled solutions were then centrifuged at 500 rpm and the concentration of phosphate anions present in the liquid phase (supernatant) was determined using UV spectrophotometer. The mass/volume ratio, used in this work, must correspond to the smallest weight of ISBM giving the highest uptake percentage of phosphate.

## 2.4 Analytical procedure and method for calculating $C_r$ ( $\text{mg}\cdot\text{L}^{-1}$ ) and $Q_{\text{ads}}$ ( $\text{mg}\cdot\text{g}^{-1}$ )

After all the kinetic and equilibrium studies, the resulting solution was filtered by Whatman filter paper ( $0.45 \mu\text{m}$ ) and the filtrate was analyzed. The phosphate anions concentrations was determined by the formation of ammonia phosphomolybdate and the subsequent reduction with ascorbic acid, followed

by spectrophotometric measurements at 700 nm, using the spectrophotometric method (AFNOR T 90-023) (RODIER, 2009).

Each experiment was carried out in duplicate and the average results are presented in this study. The variation of the adsorbed phosphate concentration ( $C_r$ ) ( $\text{mg}\cdot\text{L}^{-1}$ ) represented in the figures is defined as  $C_r = C_i - C_e$  for the solid-liquid ratio  $1 \text{ g}/40 \text{ mL}$  ( $m/V = 25 \text{ g}\cdot\text{L}^{-1}$ ). The amount of phosphate anions adsorbed at equilibrium per unit mass of biomaterial was determined according to the following equation:

$$Q_{\text{ads}} = (C_i - C_e) \times \frac{V}{m} \quad (1)$$

where  $m$  is the weight of the raw shrimp shells (g),  $V$  is the total solution volume (L),  $C_i$  and  $C_e$  are the initial and equilibrium concentration of pollutant solution ( $\text{mg}\cdot\text{L}^{-1}$ ), respectively, and  $Q_{\text{ads}}$  is the amount adsorbed at equilibrium per unit mass of ISBM ( $\text{mg}\cdot\text{g}^{-1}$ ).

The removal percentage of pollutant from aqueous solution  $X$  (%) was calculated by using the following expression:

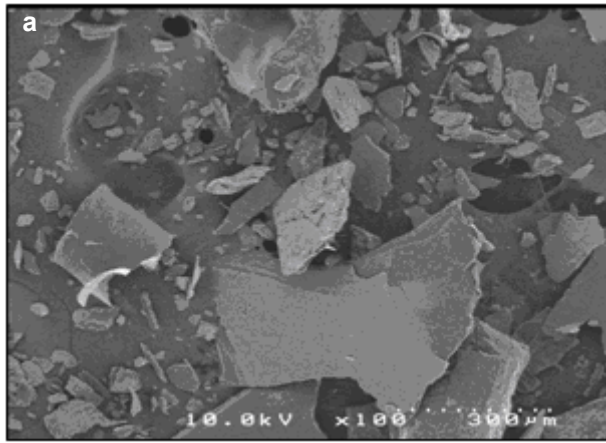
$$X(\%) = \frac{C_i - C_e}{C_i} \times 100 \quad (2)$$

# 3. RESULTS AND DISCUSSION

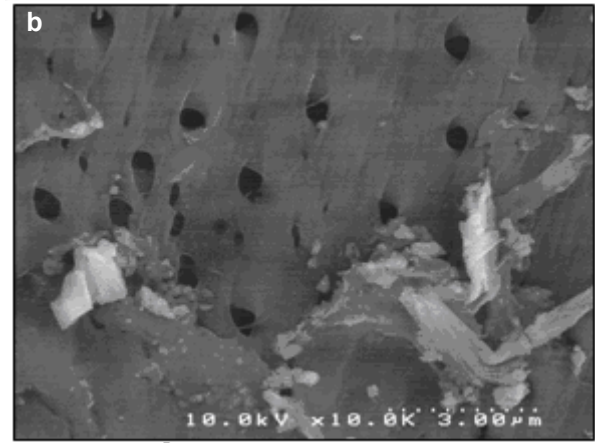
## 3.1 Characterization of the adsorbent

### 3.1.1 Analysis by scanning electron microscope

The surface micromorphology of the raw shrimp shells was investigated using scanning electron microscope (SEM, HITACHI S-4500) (Figure 2). SEM images of the raw shrimp shells microparticles (Figure 2a) shows that the dried biomaterial is not homogeneous. The SEM images indicated also the presence of grains in the structure at different sizes



Scale bar = 300 μm



Scale bar = 3.00 μm

**Figure 2.** Scanning electron microscope image of the raw shrimp shells (particle size < 500 μm).

*Image par microscopie électronique à balayage (MEB) de la poudre du broyat de carapaces de crevettes à l'état brut (granulométrie < 500 μm).*

and morphologies. Figure 2b shows the raw shrimp shells had a highly porous surface, indicating higher surface area. The morphology of this biomaterial can facilitate the adsorption of anions and metallic elements, due to the irregular surface.

### 3.1.2 Chemical analysis by X-ray emission spectroscopy

The X-ray emission spectrum of the raw shrimp shells shows that in addition to carbon and oxygen; major constituents of proteins, chitin, and carbonates, are also composed largely of calcium (Figure 3). This confirms that carbonates are mostly in the form of calcium bicarbonates as has been reported for crustacean shells (JOHNSON and PENISTON, 1982; KURITA, 2006; ROBERTS, 1992).

### 3.1.3 Infrared spectroscopy

The Infrared spectrum (IR) of the raw shrimp shells, in KBr, (Figure 4) presents typical absorption bands found in the spectrum IR of the biomolecules (proteins, lipids, glucids and nucleic acids) (VAN HOLDE *et al.*, 1998).

Hydroxyl groups (OH) and amine (NH) are represented by a broad band at 3 300  $\text{cm}^{-1}$  and absorption of these groups in the region indicates that may be an attribution to the hydrogen bonds. Stretching vibration corresponds to C-H are also common to both spectra. They are characterized by absorption in the region of 2 800  $\text{cm}^{-1}$  and 3 030  $\text{cm}^{-1}$ .

The absorption bands in the region of 1 500 and 1 650  $\text{cm}^{-1}$  corresponds to carbonyl groups stretching vibrations. The absorption bands in the region of 1 030  $\text{cm}^{-1}$  and 1 450  $\text{cm}^{-1}$  corresponds to RC-H, R-OH and amines vibration. The broad band at 1 420  $\text{cm}^{-1}$  is due to mineral salts, especially bicarbonates, characterized by CC, CO and CN groups.

## 3.2 Phosphate adsorption onto inert solid biomaterial (ISBM)

### 3.2.1 Determination of mass/volume ratio ( $R = m/V$ )

The ratio of the weight of ISBM adsorbent to volume of the aqueous phase is a very important parameter of the adsorption process. This ratio corresponding to the smallest weight of ISBM giving the highest uptake percentage of phosphate.

Different weight ( $m$ ) of ISBM were shaken with  $V = 40$  ml of phosphate solution of initial concentration value  $C_i = 100 \text{ mg}\cdot\text{L}^{-1}$  for 24 h, the temperature of 25 °C and the pH of the solutions studied is imposed by the dissolved salt for each ion. Figure 5 shows the variation of the adsorbed phosphate anions concentration ( $C_r = C_i - C_e$ ) as a function of the  $m/V$  ratio.

The results shown in Figure 5 for phosphate anions in contact with raw shell shrimp, indicated that the phosphate solution presented the highest uptake percentage up to the ratio value  $R = 25 \text{ g}\cdot\text{L}^{-1}$  and any further addition of ISBM shows no significant increasing effect on the retention process. In order to achieve experiments for maximum retention of the phosphate anions, the  $m/V$  ratio has been chosen to be  $R = 25 \text{ g}\cdot\text{L}^{-1}$  in all experiments.

### 3.2.2 Effect of contact time

The dynamics of the phosphate anions uptake was studied in batch experiments by using an ISBM suspension of 25  $\text{g}\cdot\text{L}^{-1}$  and an initial phosphate anions concentration of 50, 100, 200, 300, 500  $\text{mg}\cdot\text{L}^{-1}$  (Figure 6), the averages pH are 7.6 and 5.1 imposed by the solutes used ( $\text{NaH}_2\text{PO}_4 \cdot 12\text{H}_2\text{O}$ ) and  $\text{NaH}_2\text{PO}_4$  respectively.

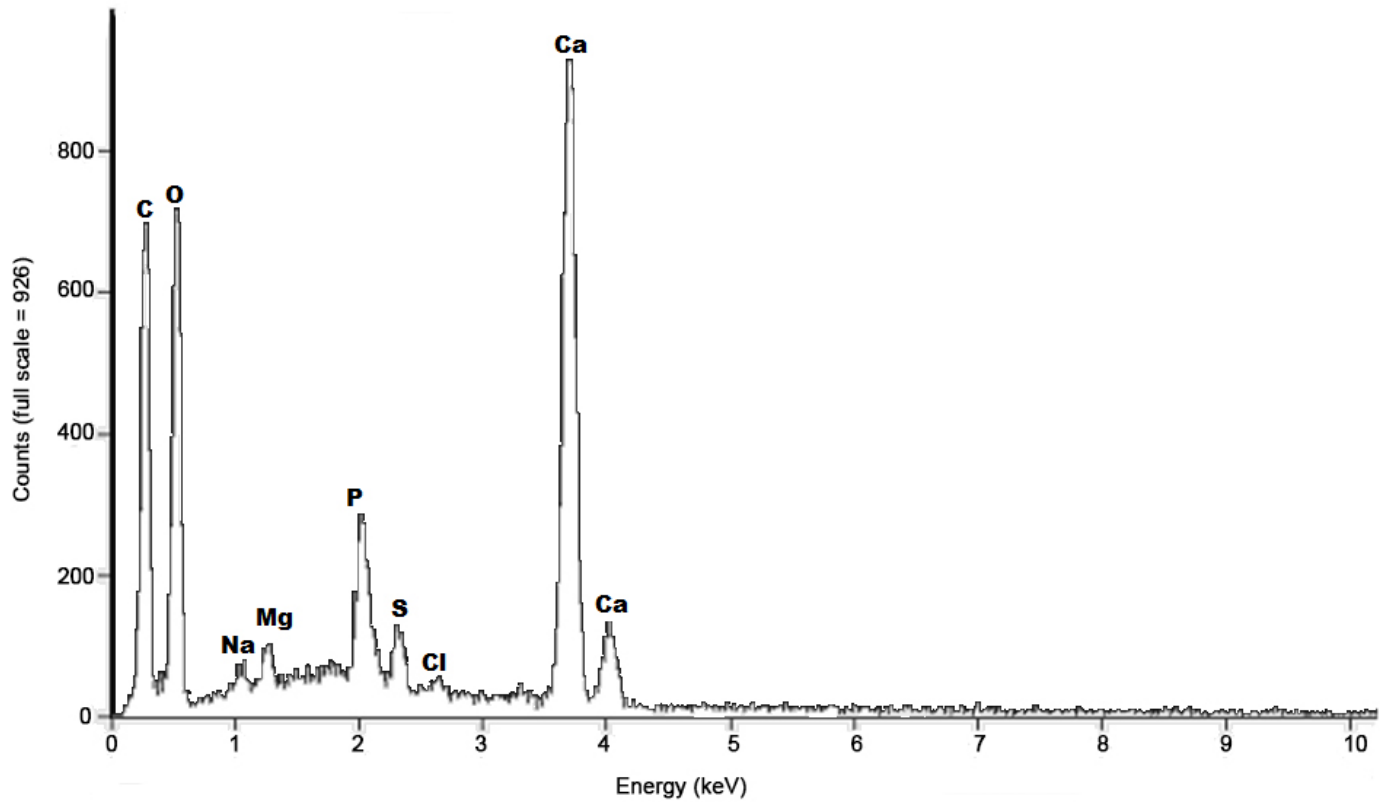


Figure 3. Energy dispersive X-ray spectrum of raw shrimp shell particles (particle size <math>< 500 \mu\text{m}</math>).  
*Spectre de rayons X à dispersion d'énergie de carapaces de crevettes à l'état brut (granulométrie <math>< 500 \mu\text{m}</math>).*

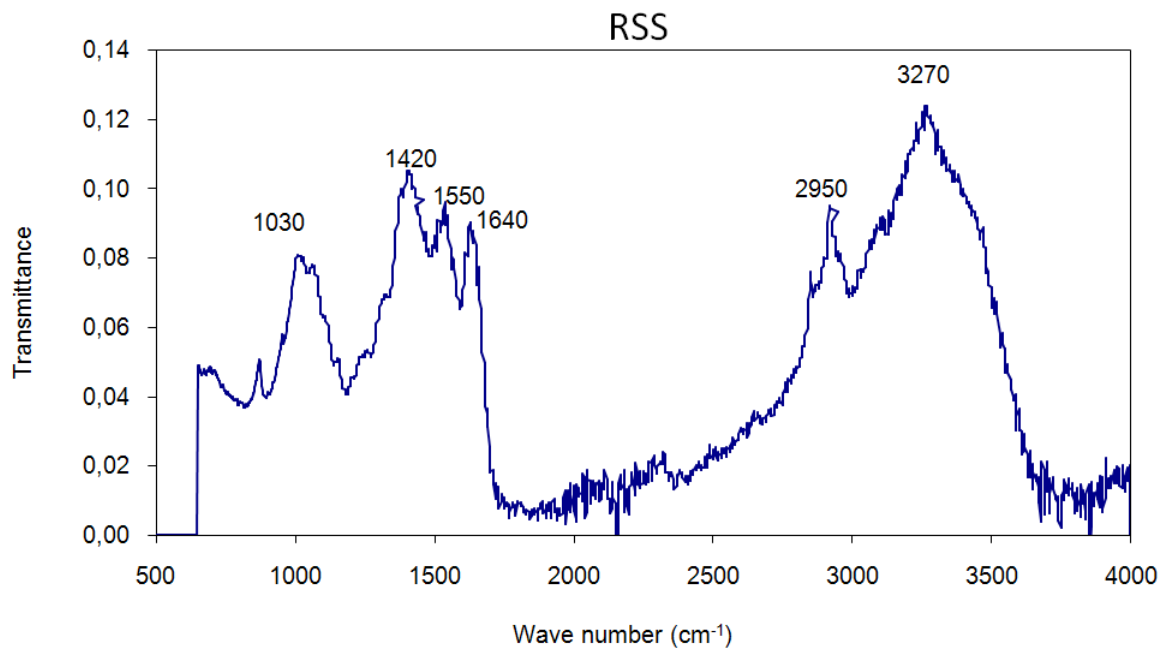


Figure 4. Infrared spectrum of the raw shrimp shells in KBr.  
*Spectre infrarouge de carapaces de crevettes à l'état brut dans KBr.*

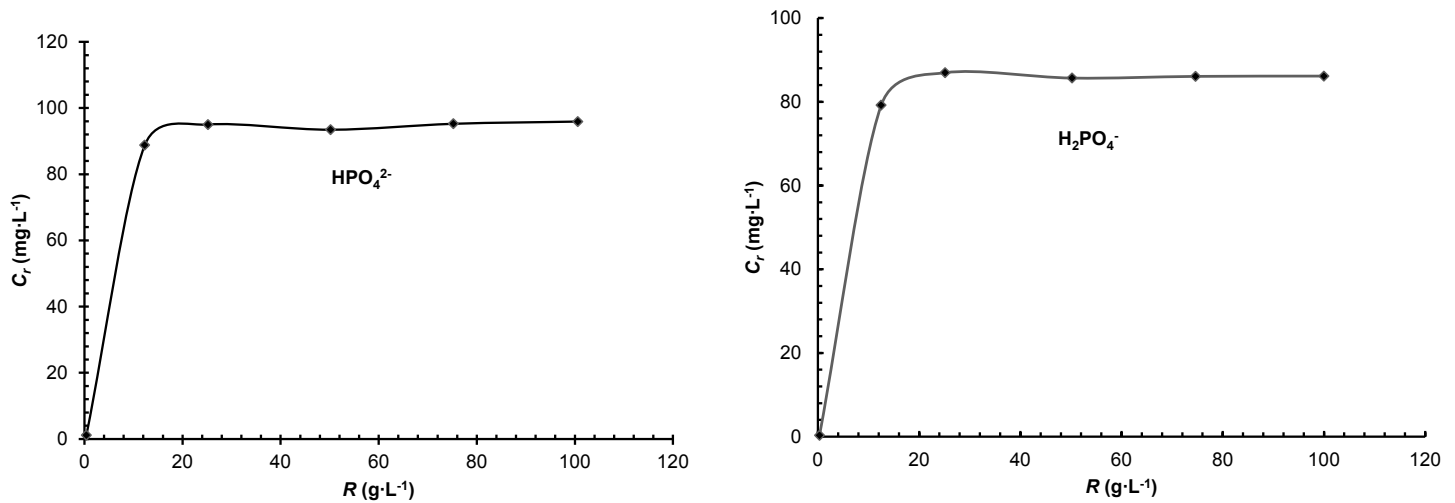


Figure 5. Variation of the phosphate anion concentrations adsorbed by the raw shrimp shells as a function of the  $m/V$  ratio ( $R$ ):  $C_i = 100 \text{ mg}\cdot\text{L}^{-1}$ ,  $T = 25 \text{ }^\circ\text{C}$  and  $t_c = 24 \text{ h}$ .

Variation de la concentration d'ions phosphates retenus par le broyat de carapaces de crevettes à l'état brut en fonction du rapport  $R$  ( $C_i = 100 \text{ mg}\cdot\text{L}^{-1}$ ;  $T = 25 \text{ }^\circ\text{C}$ ;  $t_c = 24 \text{ h}$ ).

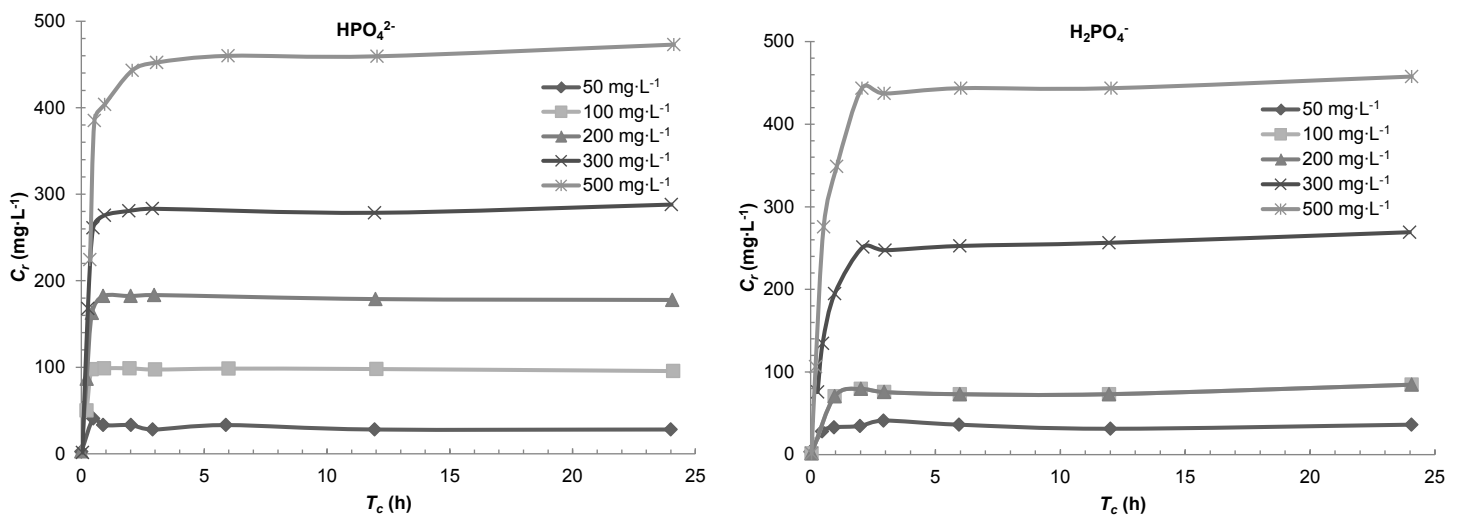


Figure 6. Effect of contact time on the removal of the phosphate anions by the raw shrimp shells ( $R = 25 \text{ g}\cdot\text{L}^{-1}$ ;  $\text{pH} = 7.55$ ;  $T = 25 \text{ }^\circ\text{C}$ ).

Variation de la concentration d'ions phosphates retenus par le broyat de carapaces de crevettes à l'état brut en fonction du temps de contact ( $R = 25 \text{ g}\cdot\text{L}^{-1}$ ;  $\text{pH} = 7.55$ ;  $T = 25 \text{ }^\circ\text{C}$ ).

The variations of the concentrations of phosphates retained versus the contact time ( $t_c$ ) are plotted in Figure 6.

The process of phosphate uptake by the ISBM appeared to follow a two phase process characterized by an initial fast retention step in less than one hour, and corresponding to an uptake concentration of about 70-80% of the initial phosphate anions concentration of the solution, followed by a much slower step before steady state. Equilibrium ( $t_c$ ) is reached in less than two hours.

Results of phosphate anions adsorption onto inert solid biomaterials of vegetal origin (SOUDANI *et al.*, 2009) showed that the adsorption variation with contact time is composed of two regimes. In the first step, the ions located in the neighborhood of the surface of the particles of the crushed biomaterial are retained and an important percentage of the phosphate anions are removed. The second step is slower than the previous one and corresponds to the rate limiting step of the adsorption process. The diffusion of the ions towards the adsorption sites buried in the ISBM particles inner structure is presumably responsible for this slow adsorption regime.

The adsorbed amounts at equilibrium of  $\text{HPO}_4^{2-}$  anions are little different than those of  $\text{H}_2\text{PO}_4^-$  anions, total ionic charge do not have a very significant effect on the adsorption process. We can therefore conclude that the adsorption of phosphate anions is a chemical adsorption and involves other interactions in addition to the electrostatic interactions. This similarity between the adsorption results of both anions is valid for low and medium concentrations below  $500 \text{ mg}\cdot\text{L}^{-1}$ , because as discussed in the following paragraph, this is not valid for high concentrations. At initial concentrations  $C_i < 500 \text{ mg}\cdot\text{L}^{-1}$  we remain well below the total saturation of adsorption sites and therefore the impact of pH is not important. This is not valid for high concentrations as explained in the following paragraph relating to the determination of maximum adsorption capacity.

To investigate the adsorption mechanism and potential speed control steps, two kinetic models were applied: the pseudo-first order Lagergren kinetic model (LAGERGREN and SVENSKA, 1889) and the pseudo-second order kinetic model (HO and MCKAY, 1998) were used. The pseudo-first order Lagergren model is expressed by:

$$\log(Q_e - Q_t) = \log(Q_e) - \frac{k_1}{2.303} t \quad (3)$$

where  $Q_e$  ( $\text{mg}\cdot\text{g}^{-1}$ ) is the adsorbed amount of phosphate ions in equilibrium,  $Q_t$  ( $\text{mg}\cdot\text{g}^{-1}$ ) adsorbed amount of phosphate ions at time  $t$  (min) and  $k_1$  ( $\text{min}^{-1}$ ) constant pseudo-first order rate.

$Q_e$  and  $k_1$  can be calculated from the slope and intercept of the  $\log(Q_e - Q_t)$  versus  $t$ . It is checked, then, if the values of the adsorption capacity in balance calculated ( $Q_{e(\text{cal})}$ ), are consistent with the experimental values of  $Q_e$  ( $Q_{e(\text{exp})}$ ) (OZACAR and SENGIL, 2003).

The kinetic model pseudo-second order is given by equation 4:

$$\frac{t}{Q_t} = \frac{1}{k_2 Q_e^2} t + Q_e \quad (4)$$

where  $k_2$  ( $\text{g}\cdot\text{mg}^{-1}\cdot\text{min}^{-1}$ ) is constant pseudo-second order rate.

The adsorption capacity  $Q_e$  at equilibrium and the constant of pseudo-second order rate  $k_2$  can be derived experimentally from the slope and intercept of the curve  $t/Q_t$  as function of  $t$ .

The parameters of kinetic model pseudo-first and pseudo-second order related to the adsorption of phosphate anions on the shrimp shells are shown in Table 1.

It is found that the adsorption of phosphate anions on shrimp shells is not a first order reaction. Although the values of the correlation coefficient ( $R^2$ ) are of the order of 0.99, the

experimental values of  $Q_{e(\text{exp})}$  are not in agreement with those calculated  $Q_{e(\text{cal})}$ .

For kinetic pseudo-second order model and as shown in Table 1, the good correlation ( $R^2 = 0.99$ ) is obtained for both initial concentrations ( $C_i = 100 \text{ mg}\cdot\text{L}^{-1}$ ,  $C_i = 300 \text{ mg}\cdot\text{L}^{-1}$  and  $C_i = 500 \text{ mg}\cdot\text{L}^{-1}$ ) and reference value  $Q_{e(\text{cal})}$  are also in good agreement with the experimental values  $Q_{e(\text{exp})}$ . This result suggests that the adsorption of phosphate anions on the shrimp shells is consistent with the pseudo-second order reaction mechanism and the adsorption rate is controlled by chemical adsorption (DOGAN *et al.*, 2004).

### 3.2.3 Determination of maximum adsorption capacity by one gram of ISBM

In order to determine the maximum adsorption capacity by one gram of ISBM derived from grinded raw shrimp shells, effect of initial concentration on the removal of  $\text{HPO}_4^{2-}$ ,  $\text{H}_2\text{PO}_4^-$  anions has been studied at varying initial concentration range between  $0.5 \text{ g}\cdot\text{L}^{-1}$  and  $7 \text{ g}\cdot\text{L}^{-1}$  (Figure 7). Initial pH range is between 7.81 and 9.28 for  $\text{HPO}_4^{2-}$  solution, and between 5.20 and 6.75 for  $\text{H}_2\text{PO}_4^-$  solution.

From these results obtained, the phosphate anions adsorbed concentrations increased with increasing initial concentration of these anions. Despite a high initial concentration equal to  $7 \text{ g}\cdot\text{L}^{-1}$  the saturation of crushed biomaterial is not reached.

In order to determine the maximum amount adsorbed, the plots  $1/C_r = f(1/C_i)$ , for both phosphate anions  $\text{HPO}_4^{2-}$  and  $\text{H}_2\text{PO}_4^-$ , allow us to evaluate the limiting adsorption concentrations ( $C_{r \text{ max}}$ ), as shown in Figure 8.

The extrapolation to infinite concentration (Figure 8), the limiting adsorption concentrations for  $\text{HPO}_4^{2-}$  and  $\text{H}_2\text{PO}_4^-$  anions corresponds to  $5.00 \text{ g}\cdot\text{L}^{-1}$  and  $10.00 \text{ g}\cdot\text{L}^{-1}$ , respectively  $0.20 \text{ g}\cdot\text{g}^{-1}$  and  $0.40 \text{ g}\cdot\text{g}^{-1}$  per unit mass of ISBM. These amounts are to much higher than those obtained with several other natural adsorbents (NGUYEN *et al.*, 2012; SOUDANI *et al.*, 2009; ISMAIL, 2012; BENYOUCEF and AMRANI, 2011; BISWAS *et al.*, 2007; KRISHNAN and HARIDAS, 2008; MEZENNER and BENSMALI, 2009). These results suggest that the raw shrimp shells have a significant potential adsorption towards the phosphate anions.

We also note that the maximum amount adsorbed for  $\text{H}_2\text{PO}_4^-$  anions is higher than this of  $\text{HPO}_4^{2-}$  anions. There is an increase in the difference between the pH of solutions of  $\text{H}_2\text{PO}_4^-$  and  $\text{HPO}_4^{2-}$  as initial concentration of these ions increases and pH of  $\text{HPO}_4^{2-}$  becomes basic. This suggests that at high concentrations of  $\text{Na}_2\text{HPO}_4$  salt, hydroxide ions compete with  $\text{HPO}_4^{2-}$  ions for the exchange sites in the inert organic matter. That can explain the increase in  $\text{HPO}_4^{2-}$  ions uptake in comparison with  $\text{H}_2\text{PO}_4^-$ . Similar behavior was observed



Table 1. Parameters of pseudo-first order and pseudo-second order kinetic models for  $\text{HPO}_4^{2-}$  and  $\text{H}_2\text{PO}_4^-$  ions.  
 Tableau 1. Paramètres des modèles cinétiques du pseudo-premier ordre et pseudo-second ordre pour les ions  $\text{HPO}_4^{2-}$  et  $\text{H}_2\text{PO}_4^-$ .

Ion	$C_i$ ( $\text{mg}\cdot\text{L}^{-1}$ )	$Q_e^{(exp)}$ ( $\text{mg}\cdot\text{g}^{-1}$ )	First-order kinetic model			Second-order kinetic model		
			$k_1$ ( $\text{min}^{-1}$ )	$Q_e^{(cal)}$ ( $\text{mg}\cdot\text{g}^{-1}$ )	$R^2$	$k_2$ ( $\text{g}\cdot\text{mg}^{-1}\cdot\text{min}^{-1}$ )	$Q_e^{(cal)}$ ( $\text{mg}\cdot\text{g}^{-1}$ )	$R^2$
$\text{H}_2\text{PO}_4^-$	100	3.38	0.0001	0.57	0.006	0.009	3.37	0.99
	300	10.76	0.0007	1.59	0.167	0.0038	10.88	0.99
	500	18.30	0.0007	1.72	0.138	0.0034	18.45	0.99
$\text{HPO}_4^{2-}$	100	3.95	0.0005	0.09	0.0078	-0.396	3.84	0.99
	300	11.52	0.002	1.44	0.206	0.556	11.50	0.99
	500	18.91	0.001	2.2	0.24	0.695	19.01	0.99

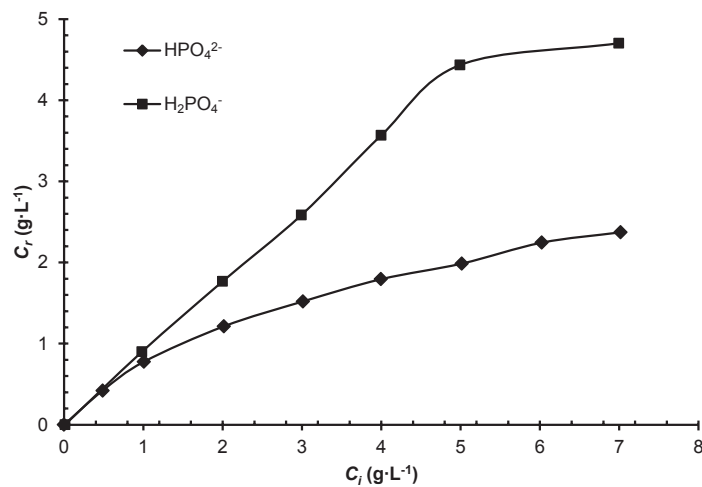


Figure 7. Effect of initial concentration on the removal of  $\text{HPO}_4^{2-}$  and  $\text{H}_2\text{PO}_4^-$  anions from single ion solutions by raw shrimp shells ( $R = 25 \text{ g}\cdot\text{L}^{-1}$ ;  $T = 25 \text{ }^\circ\text{C}$ ;  $t_c = 6 \text{ h} > t_e = 2 \text{ h}$ ).  
 Variation de la concentration d'ions phosphates retenus par le broyat de carapaces de crevettes à l'état brut en fonction de la concentration initiale ( $R = 25 \text{ g}\cdot\text{L}^{-1}$ ;  $T = 25 \text{ }^\circ\text{C}$ ;  $t_c = 6 \text{ h} > t_e = 2 \text{ h}$ ).

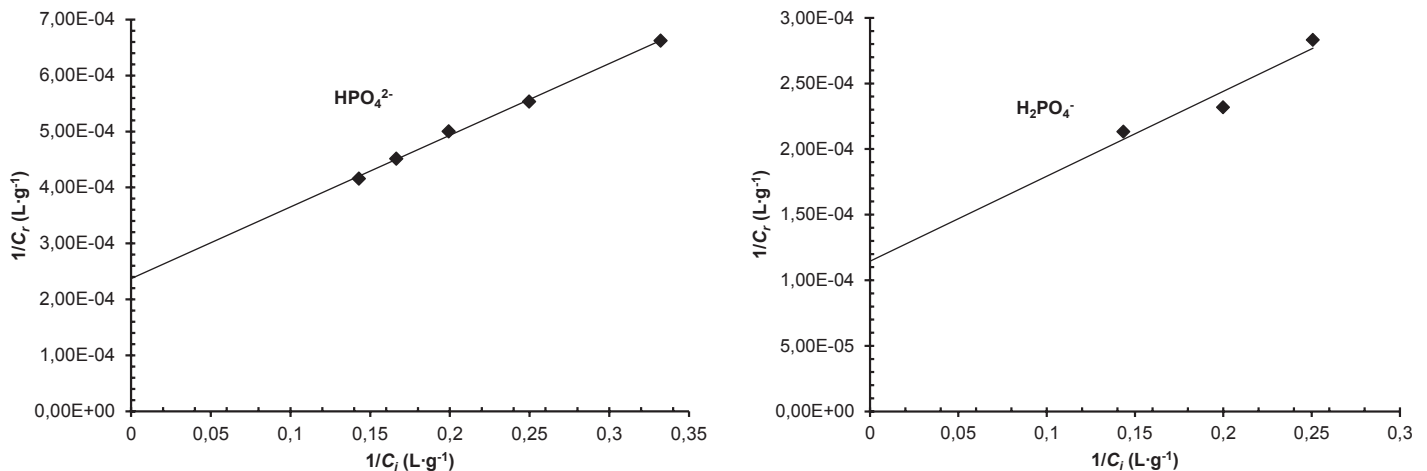


Figure 8. Plot of inverse values of  $C_r$  vs.  $C_i$  corresponding to the data shown in Figure 7.  
 Variation de  $1/C_r$  en fonction de  $1/C_i$ .

for other adsorbents of natural origin (ZHANG *et al.*, 2010; CHIBAN, 2007; SOUDANI *et al.*, 2009).

### 3.2.4 Adsorption isotherms

The adsorption isotherms are mathematical models that describe the distribution of the adsorbate species among solid and liquid phases, and are important data to understand the mechanism of the adsorption. Several models have been published in literature to describe the experimental data of the adsorption isotherms. Langmuir and Freundlich isotherms are the most frequently employed models. In this study, Langmuir, Freundlich and Temkin equations were used to describe the relationship between the adsorbed anion uptake on raw shell shrimp and its equilibrium concentration in solution. This study was carried out by varying the initial ion concentration from 0.5 g·L<sup>-1</sup> à 7 g·L<sup>-1</sup> at a temperature of 25 °C.

Langmuir isotherm (LANGMUIR, 1916):

$$\frac{1}{C_r} = \frac{1}{C_{r \max}} + \frac{1}{KC_{r \max}} \left( \frac{1}{C_e} \right) \quad (5)$$

where  $C_e$  is the equilibrium concentration (mg·L<sup>-1</sup>),  $K$  the constant related to the affinity of the binding sites and energy of adsorption (L·mg<sup>-1</sup>). A plot of  $1/C_r$  versus  $1/C_e$  should indicate a straight line of slope  $1/KC_{r \max}$  and an intercept of  $1/C_{r \max}$ . So,  $Q_{\max}$  and  $K$  can be determined.

Freundlich isotherm (FREUNDLICH, 1906):

$$\ln Q_{ads} = \ln K_f + \frac{1}{n_f} \ln C_e \quad (6)$$

where  $C_e$  is the equilibrium concentration (mg·L<sup>-1</sup>),  $Q_{ads}$  the amount of ion adsorbed (mg·g<sup>-1</sup>),  $K_f$  and  $n_f$  the Freundlich constants.

Temkin isotherm (TEMKIN and PYZHEV, 1940):

$$C_r = K \ln(C_e) + C_{te} \quad (7)$$

where  $K = f(g)$ . The plots corresponding to the Langmuir, Freundlich and Temkin isotherms for HPO<sub>4</sub><sup>2-</sup> anions are illustrated in Figure 9.

The correlation coefficients from Langmuir, Freundlich and Temkin isotherms are given in Table 2. The adsorption data in respect to HPO<sub>4</sub><sup>2-</sup> anions provide a good fit to Langmuir isotherm, based on the independence of adsorption sites and consequently the absence of interactions between the adsorbed ions. A good correlation is obtained also with the Freundlich isotherm that suggests the existence of several kinds

of adsorption sites and reinforces the good correlation obtained with the Langmuir isotherm.

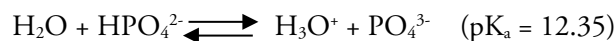
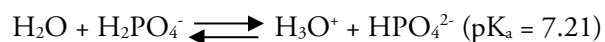
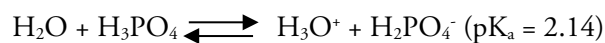
Good correlation obtained with Langmuir isotherm can be explained by the presence of several constituents inside the ISBM responsible of the phosphate adsorption, such as chitin and proteins. The factor  $g$  given by Temkin isotherm is positive, as the sign of the slope of the isotherm (logarithmic form) is positive; the interactions involved are so repulsive and weak. This confirms the good correlation of Langmuir that neglects interactions between adsorbed species.

The isotherm constant  $1/n_f$  is less than 1 ( $1/n_f < 1$ ) suggesting that the adsorption sites are not homogeneous (CHIBAN, 2007). The low value of the correlation coefficients from Temkin model of H<sub>2</sub>PO<sub>4</sub><sup>-</sup> ( $R^2 = 0.9835$ ) than HPO<sub>4</sub><sup>2-</sup> ( $R^2 = 0.9933$ ), can be related to the difference of the ionic charge of both anions.

Recall that the results of the adsorption of both anions at medium concentrations below 500 mg·L<sup>-1</sup> are slightly different (Figure 6 and Table 1), that suggest a low involvement of electrostatic interactions.

### 3.2.5 Effect of initial pH

In aqueous solution, phosphate anions exist in four forms. In strongly basic conditions, the phosphate ion (PO<sub>4</sub><sup>3-</sup>) predominates, whereas in weakly basic conditions, the hydrogen phosphate ion (HPO<sub>4</sub><sup>2-</sup>) is prevalent. In weakly acid conditions, the dihydrogen phosphate ion (H<sub>2</sub>PO<sub>4</sub><sup>-</sup>) is most common. In strongly acidic conditions, trihydrogen phosphate (H<sub>3</sub>PO<sub>4</sub>) is the main form (CHIBAN *et al.*, 2012). In aqueous solution at 25 °C are:

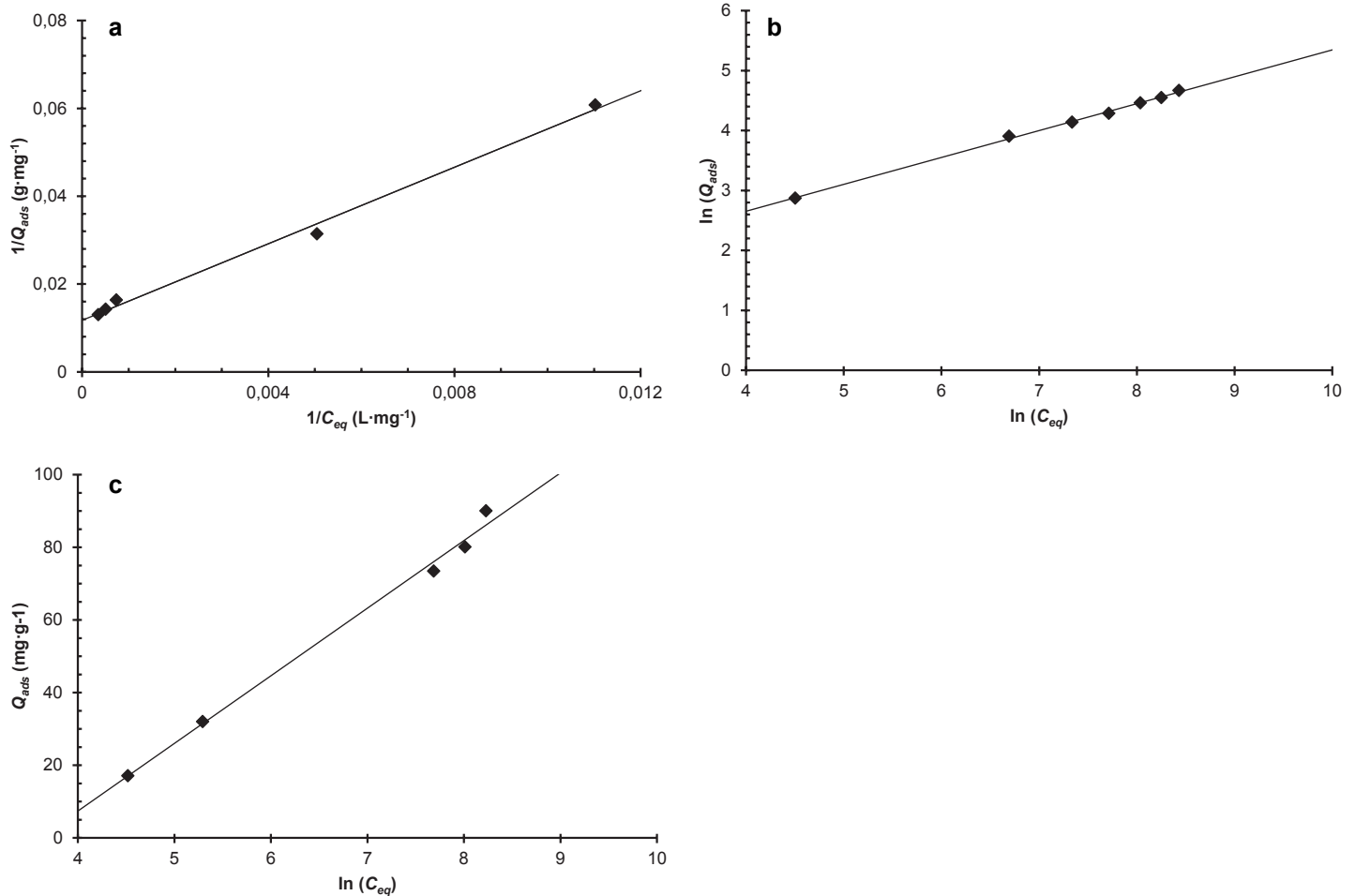


The pH of the aqueous solution is an important parameter which controls the adsorption process (KRATOCHVIL and VOLESKY, 1998). Thus, the effect of hydrogen ion concentration was examined for solutions at pH ranging from 2 to 10 at 25 °C and at fixed initial concentration ions  $C_i = 100$  mg·L<sup>-1</sup>. Figure 10 summarizes the uptake of phosphate anions at various pH by ISBM adsorbent.

It is evident that the removal of phosphate is constant at a pH range 2-6, and then decreases slightly while pH increases. The pH effect may be explained in relation to the competition effect between the hydroxide ions and mineral ions. At high

**Table 2. Isotherm model parameters and correlation coefficients for  $\text{HPO}_4^{2-}$  and  $\text{H}_2\text{PO}_4^-$  ions.**  
**Tableau 2. Paramètres et valeurs des coefficients de corrélation des isothermes pour les ions  $\text{HPO}_4^{2-}$  et  $\text{H}_2\text{PO}_4^-$ .**

Isotherm	Langmuir			Freundlich		Temkin			
	$Q_{max}$ ( $\text{g}\cdot\text{g}^{-1}$ )	$K_L$ ( $\text{L}\cdot\text{mg}^{-1}$ )	$R^2$	$1/n_f$	$K_F$ ( $\text{mg}^{1-1/n}\cdot\text{L}^{1/n}\cdot\text{g}^{-1}$ )	$R^2$	$K_T$ ( $\text{L}\cdot\text{g}^{-1}$ )	$g$	$R^2$
$\text{HPO}_4^{2-}$	0.092	0.0024	0.99	0.46	2.12	0.99	462.95	>0	0.993
$\text{H}_2\text{PO}_4^-$	0.133	0.0062	0.99	0.45	5.58	0.98	643.12	>0	0.983



**Figure 9. Isotherm plots for the adsorption of  $\text{HPO}_4^{2-}$  anions by the raw shrimp shells: a) Langmuir isotherm; b) Freundlich isotherm; c) Temkin isotherm.**

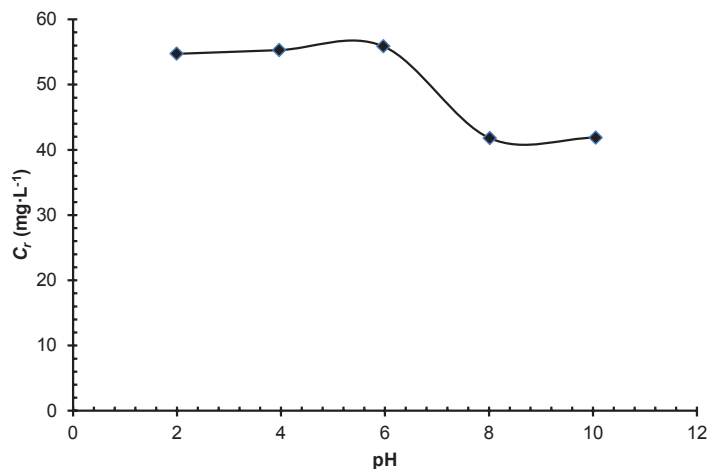
*Isothermes d'adsorption de  $\text{HPO}_4^{2-}$  sur le broyat de carapaces de crevettes à l'état brut : a) isotherme de Langmuir; b) isotherme de Freundlich; c) isotherme de Temkin.*

pH values, the concentration of  $\text{OH}^-$  far exceeds that of the mineral ions; hence these are bound to the adsorbent, leaving the mineral ions unbound as was explained in the paragraph 3.2.2 (CHIBAN, 2007; SOUDANI *et al.*, 2009; ZHANG *et al.*, 2010).

The results of the zero point of charge of the shrimp shells is found to be  $\text{pH}_{\text{PZC}} = 8.1$ . This shows that at pH less than 8.1 the surface of the shrimp shells is predominated by

positive charges while at pH greater than 8.1 the surface is predominated by negative charges. Thus, at  $\text{pH} < \text{pH}_{\text{PZC}}$ , the surface has a high positive charge density; uptake of negatively charged phosphate ions would be high. At  $\text{pH} > \text{pH}_{\text{PZC}}$ , the surface has a high negative charge density; uptake of negatively charged phosphate ions would be low.

The reason for good removal of orthophosphates at the lower pH is that the negative charge on the surface is reduced



**Figure 10.** Effect of pH on the removal of phosphate anions by the raw shrimp shells ( $C_i = 100 \text{ mg}\cdot\text{L}^{-1}$ ,  $R = 25 \text{ g}\cdot\text{L}^{-1}$ ,  $t_c = 24 \text{ h}$ ,  $T = 25 \text{ }^\circ\text{C}$ ).

*Variation de la concentration d'ions phosphates retenus par le broyat de carapaces de crevettes à l'état brut en fonction du pH ( $C_i = 100 \text{ mg}\cdot\text{L}^{-1}$ ,  $R = 25 \text{ g}\cdot\text{L}^{-1}$ ,  $t_c = 24 \text{ h}$ ,  $T = 25 \text{ }^\circ\text{C}$ ).*

due to the excess of protons in solutions. As a result, the pH of the system decreases and the number of positively charged sites increase. A positively charged surface site on the raw shell shrimp favors the adsorption of the phosphate anions due to electrostatic attraction.

### 3.2.6 Effect of temperature

The effect of temperature ( $T$ ) on the phosphate anions adsorption by the raw shrimp shells were studied at 20, 25, 30, 35, and 40 °C at 100 mg·L<sup>-1</sup> initial anions concentration and at natural pH. The results obtained, were plotted in Figure 11 for  $\text{HPO}_4^{2-}$ ,  $\text{H}_2\text{PO}_4^-$  anions.

These figures show that the concentration of phosphate anions adsorbed at different temperatures indicated the removal of these anions decreases with increasing of the temperature to stabilize from  $T = 30 \text{ }^\circ\text{C}$ . These results indicated also that the adsorption process of anions studied was exothermic. The adsorption capacity decreased with increasing of temperature. At  $T > 30 \text{ }^\circ\text{C}$ , the influence of temperature on the adsorption becomes negligible.

### 3.3 Comparison with published data

The application of low-cost and easily available materials in wastewater treatment has been widely investigated during recent years. Particularly, the phosphate adsorption on different materials has been widely studied during recent years. In order to situate our natural adsorbent among those used to remove orthophosphate from aqueous solutions, the

orthophosphate anions adsorption capacity of Shrimp shells adsorbent was compared to the adsorption capacities of some other adsorbents reported in literature (Table 3). For example, NAMASIVAYAM and SANGEETHA (2004) examined the adsorption of phosphate onto coir pith carbon and they have found a 5.1 mg·g<sup>-1</sup> phosphate uptake capacity. OGUZ (2004) studied for the blast furnace slag the adsorption of phosphate ions and the adsorption capacity of phosphate was only 6.37 mg·g<sup>-1</sup> at an equilibrium time ( $t_e$ ) of 60 min and a pH of 8.5. HUANG et al. (2008) studied the phosphate removal from wastewater using red mud. They found that all activated red mud samples show higher surface area and total pore volume as well as higher adsorption capacity for phosphate removal. The red mud with HCl treatment shows the highest adsorption capacity among all the red mud samples, giving adsorption capacity of 0.58 mg·g<sup>-1</sup> at pH 5.5 and 40 °C. Peat was also used by XIONG and MAHMOOD (2010) as an adsorbent without any pretreatment for studying the adsorption of phosphate from aqueous solution. They found that the maximum phosphate adsorption on peat was 8.91 mg·g<sup>-1</sup>. It is obvious from these results that the adsorption affinity of Shrimp shells adsorbent towards phosphate ions is comparable or higher than other available adsorbents. The differences in phosphate anions maximum adsorption capacities are due to the properties of each adsorbent such as structure, surface area, polyphenolic groups, functional groups etc. It is proved that Shrimp shells could be considered as a promising material and one alternative source for low cost adsorbents to remove phosphate ions from wastewaters.

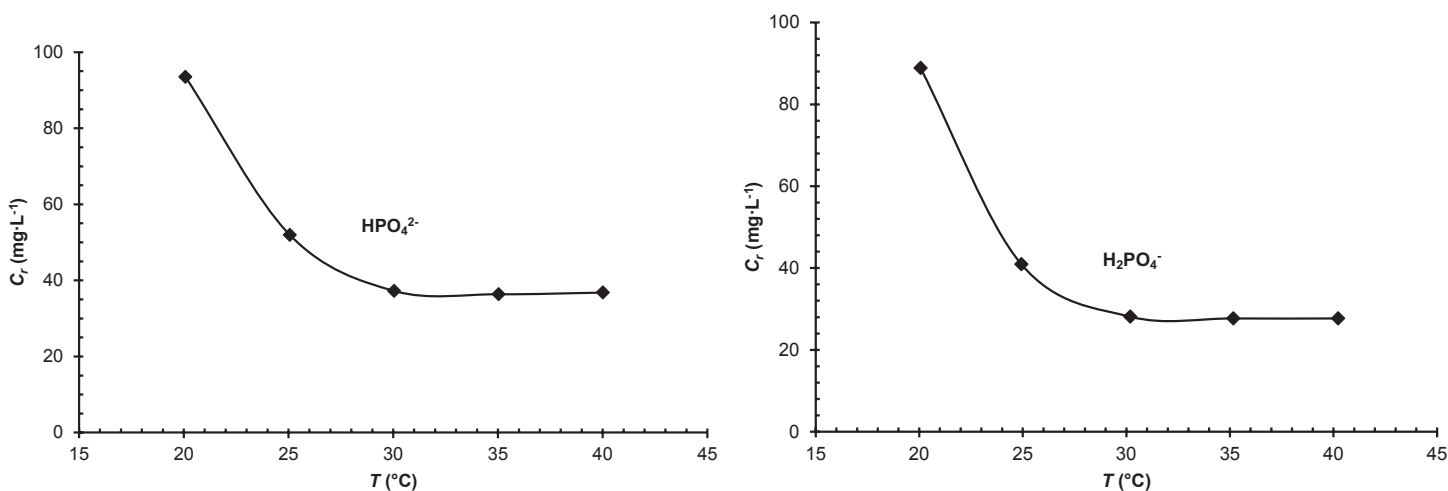
## 4. CONCLUSION

The results of this study indicate that ISBM is an effective adsorbent for removal of phosphate from aqueous solutions. At saturation, the adsorption capacities of phosphate ions by 1 gram of ISBM was found to be about 200 mg·g<sup>-1</sup> and 400 mg·g<sup>-1</sup> for  $\text{HPO}_4^{2-}$  and  $\text{H}_2\text{PO}_4^-$  respectively. It was found that the concentration of phosphate adsorbed on ISBM comparatively depends on pH, temperature and contact time of the solution. The removal of these anions decreases with increasing of the temperature to stabilize from  $T = 30 \text{ }^\circ\text{C}$  and the adsorption process was exothermic. Results indicated also that the removal of phosphate is constant at a pH range 2-6, and then decreases slightly while pH increases. Langmuir and Freundlich isotherms show the existence of several independent adsorption sites obeying to Langmuir. The Temkin isotherm suggests repulsive and weak electrostatic interactions.

**Table 3.** Comparison of the phosphate adsorption capacities onto shrimp shells with other natural adsorbents.

**Tableau 3.** Comparaison des quantités maximales d'ions orthophosphates adsorbées par différents matériaux.

Adsorbent	$Q_{max}$ ( $\text{mg}\cdot\text{g}^{-1}$ )	References
A mixture of sand and dolomite	0.17	PROCHASKA and ZOUBOULIS (2006)
Activated coir pith carbon	5.1	NAMASIVAYAM and SANGEETHA (2004)
Active red mud	31	LIU <i>et al.</i> (2007)
Apatite	4.76	MOLLE <i>et al.</i> (2005)
Bauxite	2.95	ALTUNDOGAN and TUMEN (2002)
Blast furnace slag	6.37	OGUZ (2004)
Calcite	1.9	KARAGEORGIU <i>et al.</i> (2007)
<i>Carpobrotus edulis</i>	25.6	CHIBAN (2007)
Date palm fibers	5.85	RIAHI <i>et al.</i> (2013)
Fly ash	6.6	AGYEI <i>et al.</i> (2002)
Iron oxide tailings	8	ZENG <i>et al.</i> (2004)
<i>Launea arborescens</i>	17.7	CHIBAN (2007)
Modified palygorskites	9	YE <i>et al.</i> (2006)
Naturel palygorskites	4	YE <i>et al.</i> (2006)
Peat	8.91	XIONG and MAHMOOD (2010)
<i>Posidonia oceanica</i> (L.) fibers	7.45	WAHAB <i>et al.</i> (2011)
Red mud	0.58	HUANG <i>et al.</i> (2008)
Sands	0.42	DEL BUBBA <i>et al.</i> (2003)
<i>Senecio anthophorbium</i>	24.42	CHIBAN (2007)
<i>Withania frutescens</i>	23.2	CHIBAN (2007)
Wood fiber treated	4.3	EBERHARDT <i>et al.</i> (2006)
Zeolite	2.15	SAKADEVAN and BAVOR (1998)
Shrimp shells	200	This study



**Figure 11.** Effect of temperature on the removal of phosphate anions by the raw shrimp shells ( $R = 25 \text{ g}\cdot\text{L}^{-1}$ ;  $t_c = 24 \text{ h}$ ;  $C_i = 100 \text{ mg}\cdot\text{L}^{-1}$ ).

Variation de la concentration d'ions phosphates retenus par le broyat de carapaces de crevettes à l'état brut en fonction de la température ( $R = 25 \text{ g}\cdot\text{L}^{-1}$ ;  $t_c = 24 \text{ h}$ ;  $C_i = 100 \text{ mg}\cdot\text{L}^{-1}$ ).

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